

# BIOLOGY MDCAT

## UNIT-5 (A+ Series)

### TOPIC:

#### ✓ Diversity among Animals

- Q.1 All animals are:  
A. Photoautotrophs  
B. Absorptive heterotrophs  
C. Ingestive heterotrophs  
D. Chemoautotrophs
- Q.2 Animals of Phylum \_\_\_\_\_ lack symmetry.  
A. Porifera  
B. Platyhelminthes  
C. Cnidaria  
D. Nematoda
- Q.3 The body of tapeworm is ribbon-like and divided into segments called:  
A. Proglottids  
B. Glottids  
C. Stomium  
D. Prostomium
- Q.4 In which of the following organism, similar body parts are arranged as spokes or radiate from the central body axis?  
A. *Euplectella*  
B. Sea anemone  
C. *Spongilla*  
D. *Sycon*
- Q.5 Which of the following is involved in digestion in sponges?  
A. Mesenchyme  
B. Choanocytes  
C. Pinacocytes  
D. Spongocoel
- Q.6 All of the following are the characteristics of protostomes except:  
A. Schizocoelom  
B. Bilateral symmetry  
C. Blastopore form mouth  
D. Radial cleavage
- Q.7 Largest living vertebrate belongs to:  
A. Amphibians  
B. Aves  
C. Reptiles  
D. Mammals
- Q.8 Which of the following phyla has largest invertebrate?  
A. Arthropoda  
B. Mollusca  
C. Cnidaria  
D. Nematoda
- Q.9 Lining of the digestive tract and associated glands are derived from:  
A. Ectoderm  
B. Endoderm  
C. Mesoderm  
D. Blastoderm
- Q.10 Which of the following has a gastrovascular cavity?  
A. Sponges  
B. Roundworms  
C. Earthworms  
D. Flatworms
- Q.11 Sea cucumber belongs to Phylum:  
A. Cnidaria  
B. Platyhelminthes  
C. Echinodermata  
D. Nematoda
- Q.12 Coral reef formation is characteristic of:  
A. Cnidarians  
B. Sponges  
C. Annelids  
D. Molluscs
- Q.13 Offensive and defensive structures of cnidarians are:  
A. Flagella  
B. Collar cells  
C. Nematocysts  
D. Spicules
- Q.14 Which of the following system is well-developed in parasitic flatworms?  
A. Respiratory system  
B. Reproductive system  
C. Digestive system  
D. Muscular system
- Q.15 First time tube-like digestive system was developed in:  
A. *Actinia*  
B. *Ascaris*  
C. *Neries*  
D. *Chaetopterus*
- Q.16 Tracheal system for respiration is found in phylum \_\_\_\_\_.  
A. Arthropoda  
B. Mollusca  
C. Annelida  
D. Echinodermata

- Q.17 Water vascular system is characteristic of:  
 A. *Asterias* B. *Aurelia*  
 C. *Limax* D. *Dugesia*
- Q.18 The one which lacks a digestive system:  
 A. *Hydra* B. *Ascaris*  
 C. *Planaria* D. *Taenia*
- Q.19 A structure which is not helpful for locomotion in annelids is:  
 A. Jointed legs B. Setae  
 C. Hydrostatic skeleton D. Longitudinal and circular muscles
- Q.20 Flame cell system is found in:  
 A. Flatworms B. Hemichordates  
 C. Roundworms D. Echinoderms
- Q.21 Excretory system of \_\_\_\_\_ is associated with digestive tract.  
 A. Insects B. *Planaria*  
 C. Earthworms D. Tapeworms
- Q.22 The cavity between body wall and alimentary canal in *Rhabditis* is known as:  
 A. Coelom B. Enteron  
 C. Pseudocoelom D. Spongocoel
- Q.23 The first opening of the embryo in all animals is:  
 A. Mouth B. Anus  
 C. Blastopore D. Archenteron
- Q.24 The shell is much reduced and internal in:  
 A. Garden snail B. *Anodonta*  
 C. Oyster D. *Sepia*
- Q.25 It is an example of protostomes:  
 A. Sponges B. Segmented worms  
 C. Echinoderms D. Cnidarians
- Q.26 Which of the following animal is involved in the churning of soil and commonly termed a natural plough?  
 A. Leech B. Earthworm  
 C. Neries D. Roundworm
- Q.27 It is commonly known as hookworm:  
 A. *Ancylostoma duodenale* B. *Ascaris lumbricoides*  
 C. *Enterobius vermicularis* D. *Pheritima posthuma*
- Q.28 In animals, reproductive system develops from:  
 A. Mesenchyme B. Ectoderm  
 C. Endoderm D. Mesoderm
- Q.29 Arthropods are believed to have a common origin with annelids because both have some common characteristics such as:  
 A. Segment body B. Cuticle  
 C. Appendages D. All A, B, C
- Q.30 Aquatic arthropods respire through:  
 A. Gills only B. Gills and book lungs  
 C. Tracheoles D. Skin
- Q.31 Spider belongs to which class of phylum Arthropoda?  
 A. Insecta B. Crustacea  
 C. Myriapoda D. Arachnida
- Q.32 Bony fishes contain \_\_\_\_\_ cranial nerves.  
 A. 5 B. 15  
 C. 10 D. 20
- Q.33 Amphibians are thought to be evolved from:  
 A. Cartilaginous fishes B. Hag fishes  
 C. Dipnoi fishes D. Bony fishes

- Q.34 It is not a characteristic of chondrichthyes:  
 A. Body fusiform  
 B. Ctenoid scales  
 C. Mouth is usually ventral  
 D. Gills are not covered by an operculum
- Q.35 Which of the following animal lives in our large intestine and its movement causes intense itching of anus?  
 A. *Enterobius vermicularis*  
 B. *Ascaris lumbricoides*  
 C. *Ancylostoma duodenale*  
 D. *Pheritima posthuma*
- Q.36 *Sea anemone* belong to Phylum:  
 A. Coelenterata  
 B. Arthropoda  
 C. Echinodermata  
 D. Annelida
- Q.37 A connecting link between birds and reptiles is:  
 A. *Archaeopteryx*  
 B. Varnope  
 C. Cotylosaurs  
 D. Lobed fins fishes
- Q.38 A reptile has a completely partitioned four-chambered heart is:  
 A. Camel  
 B. Lizard  
 C. Crocodile  
 D. Snake
- Q.39 Most developed mammals belong to:  
 A. Prototheria  
 B. Eutheria  
 C. Metatheria  
 D. Marsupial mammals
- Q.40 First time copulatory organ was developed in:  
 A. Fishes  
 B. Reptiles  
 C. Amphibians  
 D. Birds
- Q.41 Shark belongs to:  
 A. Cartilaginous fishes  
 B. Agnatha  
 C. Bony fishes  
 D. Cyclostomata
- Q.42 Opossum and Tasmanian wolf belong to:  
 A. Prototheria  
 B. Eutheria  
 C. Metatheria  
 D. Monotremes
- Q.43 *Vibrio cholerae* can be transmitted through:  
 A. Mosquito  
 B. Tsetse fly  
 C. House fly  
 D. Locust
- Q.44 Which of the following parasite produces anticoagulant during feeding?  
 A. Pinworm  
 B. Liver fluke  
 C. Hookworm  
 D. *Rhabditis*
- Q.45 Which one is not a parasite?  
 A. Planaria  
 B. *Trypanosoma*  
 C. *Plasmodium*  
 D. Tapeworm
- Q.46 Alternation of generation and polymorphism is characteristic of:  
 A. Cnidarians  
 B. Sponges  
 C. Round worms  
 D. Segmented worms
- Q.47 Sponges are thought to have evolved from:  
 A. Choanoflagellates  
 B. Actinopods  
 C. Ciliates  
 D. Foraminiferans
- Q.48 Collar cells of sponges are:  
 A. Flagellated  
 B. Ciliated  
 C. Non-flagellated  
 D. Pinacocytes
- Q.49 Jawless fishes belong to:  
 A. Agnatha  
 B. Chondrichthyes  
 C. Gnathostomata  
 D. Osteichthyes
- Q.50 All of the following are poikilotherms except:  
 A. Amphibians  
 B. Reptiles  
 C. Pisces  
 D. Birds
- Q.51 Right aortic arch is characteristic of:  
 A. Mammals  
 B. Reptiles  
 C. Birds  
 D. Amphibians

- Q.52 Echinoderms and chordates have:  
 A. Pseudocoelom  
 B. Enterocoelom  
 C. Schizocoelom  
 D. Hemocoel
- Q.53 *Enterobius vermicularis* is commonly called:  
 A. Shipworms  
 B. Hook worms  
 C. Bookworms  
 D. Pinworms
- Q.54 Locomotory organs of *L. terrestris* are:  
 A. Flagella  
 B. Parapodia  
 C. Setae  
 D. Cilia
- Q.55 Colonial corals secrete  $\text{CaCO}_3$  from:  
 A. Cnidocytes  
 B. Epidermal cells  
 C. Nematocyst  
 D. Endodermal cells
- Q.56 Which of the following is not a body region of *Saccoglossus*?  
 A. Collar  
 B. Mouth  
 C. Proboscis  
 D. Trunk
- Q.57 Voice producing organ in birds is:  
 A. Larynx  
 B. Pharynx  
 C. Syrinx  
 D. Tongue
- Q.58 *Ascaris* normally inhabits the lumen of:  
 A. Stomach  
 B. Colon  
 C. Small intestine  
 D. Rectum
- Q.59 A freshwater coelenterate that lacks alternation of generation:  
 A. Hydra  
 B. *Aurelia*  
 C. *Obelia*  
 D. *Actinia*
- Q.60 Which of the following are spicules and spongin fibers forming cells?  
 A. Pinacocytes  
 B. Choanocytes  
 C. Scleroblasts  
 D. Neoblasts

Mdca Site

| A  | B | C | D | A | B  | C | D | A | B | C  | D | A | B | C | D  |   |   |   |   |
|----|---|---|---|---|----|---|---|---|---|----|---|---|---|---|----|---|---|---|---|
| 1  | ○ | ○ | ● | ○ | 16 | ● | ○ | ○ | ○ | 31 | ○ | ○ | ○ | ● | 46 | ● | ○ | ○ | ○ |
| 2  | ● | ○ | ○ | ○ | 17 | ● | ○ | ○ | ○ | 32 | ○ | ○ | ○ | ○ | 47 | ● | ○ | ○ | ○ |
| 3  | ● | ○ | ○ | ○ | 18 | ● | ○ | ○ | ○ | 33 | ○ | ○ | ○ | ○ | 48 | ● | ○ | ○ | ○ |
| 4  | ○ | ○ | ○ | ○ | 19 | ● | ○ | ○ | ○ | 34 | ○ | ○ | ○ | ○ | 49 | ● | ○ | ○ | ○ |
| 5  | ○ | ○ | ○ | ○ | 20 | ● | ○ | ○ | ○ | 35 | ○ | ○ | ○ | ○ | 50 | ○ | ○ | ○ | ○ |
| 6  | ○ | ○ | ○ | ○ | 21 | ● | ○ | ○ | ○ | 36 | ○ | ○ | ○ | ○ | 51 | ○ | ○ | ○ | ○ |
| 7  | ○ | ○ | ○ | ○ | 22 | ○ | ○ | ○ | ○ | 37 | ○ | ○ | ○ | ○ | 52 | ○ | ○ | ○ | ○ |
| 8  | ○ | ○ | ○ | ○ | 23 | ○ | ○ | ○ | ○ | 38 | ○ | ○ | ○ | ○ | 53 | ○ | ○ | ○ | ○ |
| 9  | ○ | ○ | ○ | ○ | 24 | ○ | ○ | ○ | ○ | 39 | ○ | ○ | ○ | ○ | 54 | ○ | ○ | ○ | ○ |
| 10 | ○ | ○ | ○ | ○ | 25 | ○ | ○ | ○ | ○ | 40 | ○ | ○ | ○ | ○ | 55 | ○ | ○ | ○ | ○ |
| 11 | ○ | ○ | ○ | ○ | 26 | ○ | ○ | ○ | ○ | 41 | ○ | ○ | ○ | ○ | 56 | ○ | ○ | ○ | ○ |
| 12 | ○ | ○ | ○ | ○ | 27 | ○ | ○ | ○ | ○ | 42 | ○ | ○ | ○ | ○ | 57 | ○ | ○ | ○ | ○ |
| 13 | ○ | ○ | ○ | ○ | 28 | ○ | ○ | ○ | ○ | 43 | ○ | ○ | ○ | ○ | 58 | ○ | ○ | ○ | ○ |
| 14 | ○ | ○ | ○ | ○ | 29 | ○ | ○ | ○ | ○ | 44 | ○ | ○ | ○ | ○ | 59 | ○ | ○ | ○ | ○ |
| 15 | ○ | ○ | ○ | ○ | 30 | ○ | ○ | ○ | ○ | 45 | ○ | ○ | ○ | ○ | 60 | ○ | ○ | ○ | ○ |

| Roll No. |   |   |   |   |   |   |
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|          |   |   |   | 2 | 6 | 5 |
| 0        | 0 | 0 | 0 | 0 | 0 | 0 |
| 1        | 1 | 1 | 1 | 1 | 1 | 1 |
| 2        | 2 | 2 | 2 | 2 | 2 | 2 |
| 3        | 3 | 3 | 3 | 3 | 3 | 3 |
| 4        | 4 | 4 | 4 | 4 | 4 | 4 |
| 5        | 5 | 5 | 5 | 5 | 5 | 5 |
| 6        | 6 | 6 | 6 | 6 | 6 | 6 |
| 7        | 7 | 7 | 7 | 7 | 7 | 7 |
| 8        | 8 | 8 | 8 | 8 | 8 | 8 |
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SUBJECT: Bio

NAME: M. L.

R. NO. (IN WORDS):

CLASS: 8-4

SIGN: Hay

# CHEMISTRY MDCAT

## UNIT-5 (A + SERIES)

TOPIC:-

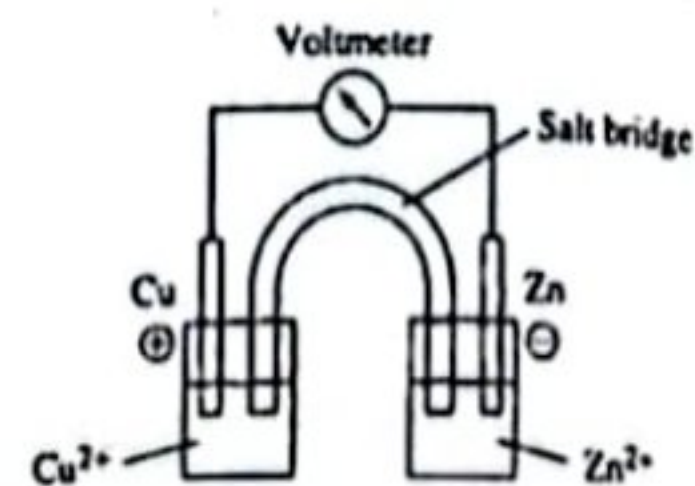
- ✓ THERMOCHEMISTRY AND ENERGETICS OF CHEMICAL REACTIONS
- ✓ ELECTROCHEMISTRY

- Q.1 The oxidation states of the most electronegative element in the reaction of  $\text{BaO}_2$  with dil.  $\text{H}_2\text{SO}_4$  are:  
A. 0 and -1  
B. -1 and -2  
C. -2 and 0  
D. -2 and +1
- Q.2 The oxidation number of Fe in  $\text{K}_4[\text{Fe}(\text{CN})_6]$  is  
A. +6  
B. +4  
C. +3  
D. +2
- Q.3 The oxidation number of oxygen in potassium superoxide  $\text{KO}_2$   
A. -2  
B. -1  
C. -1/2  
D. -1/4
- Q.4 The reaction  $\text{H}_2\text{S} + \text{H}_2\text{O}_2 \longrightarrow \text{S} + 2\text{H}_2\text{O}$  shows:  
A. acidic nature of  $\text{H}_2\text{O}_2$   
B. alkaline nature of  $\text{H}_2\text{O}_2$   
C. Oxidizing action of  $\text{H}_2\text{O}_2$   
D. reducing action of  $\text{H}_2\text{O}_2$
- Q.5 Electrolysis of a copper sulfate solution yields at inert electrodes:  
A. Cu,  $\text{SO}_2$   
B. Cu,  $\text{H}_2$   
C. Cu,  $\text{O}_2$   
D.  $\text{H}_2$ ,  $\text{O}_2$
- Q.6 When two moles of  $\text{H}_2(\text{g})$  and one mole of  $\text{O}_2(\text{g})$  react to form  $\text{H}_2\text{O}(\text{g})$  484KJ heat is evolved the  $\Delta H^\circ_f$  of  $\text{H}_2\text{O}(\text{g})$   
A.  $-286\text{KJmol}^{-1}$   
B.  $-242\text{KJmol}^{-1}$   
C.  $-484\text{KJmol}^{-1}$   
D.  $+242\text{KJmol}^{-1}$
- Q.7 Which is not a state function?  
A. Heat contents  
B. Heat  
C. Internal energy  
D. Pressure
- Q.8 The heat absorbed in a reaction at constant temperature and volume is equal to:  
A.  $\Delta E$   
B.  $\Delta H$   
C.  $-\Delta E$   
D.  $-\Delta H$
- Q.9 Given the following data  $\text{C} + \text{D} \rightarrow \text{A} + \text{B}$   $\Delta H^\circ = 48 \text{ kJ}$ ,  
 $\text{C} + \text{D} \rightarrow \text{E}$ ,  $\Delta H = 72 \text{ KJ}$ . The value of  $\Delta H^\circ$  (kJ) for the reaction  $\text{A} + \text{B} \rightarrow \text{E}$   
A. 120  
B. 24  
C. 60  
D. -24
- Q.10  $2\text{CH}_3\text{OH}(\text{l}) + 3\text{O}_2(\text{g}) \longrightarrow 2\text{CO}_2(\text{g}) + 4\text{H}_2\text{O}(\text{g})$  two mole of methanol when burnt in oxygen gives out  $1446 \text{ KJ mol}^{-1}$  heat. If one mole of oxygen is used, the amount of heat evolved:  
A. 723 KJ  
B. 964 KJ  
C. 241 KJ  
D. 482 KJ
- Q.11 On electrolysis of dil.  $\text{H}_2\text{SO}_4$  using platinum electrodes, the product obtained at the anode will be  
A. Hydrogen  
B. Oxygen  
C. Sulphur Dioxide  
D. Hydrogen Sulphide
- Q.12 Warming ammonium chloride with sodium hydroxide in a test tube is an example of  
A. Closed System  
B. Open System  
C. Isolated System  
D. None of these
- Q.13 If 32g of methane is burned in excess of oxygen then standard enthalpy of combustion is  
A. Increased two times  
B. Reduced to half  
C. Increase four times  
D. Reduced to four times
- Q.14 The enthalpy change which is exothermic in Born-Haber's Cycle  
A.  $\Delta H_{\text{IE}}$   
B.  $\Delta H_{\text{EA}}$   
C.  $\Delta H_{\text{at}}$   
D.  $\Delta H_{\text{diss}}$



Q.15 In the cell shown below, which of the following is (are) true?

- I. Electrons flow through the meter from left to right
- II. Cu is the anode
- III. The spontaneous reaction is  $\text{Cu}^{2+} + \text{Zn} \rightarrow \text{Cu} + \text{Zn}^{2+}$



- A. I only
- B. III only
- C. II only
- D. I and II

Q.16 The net heat change in a chemical reaction is same whether it is brought in two or more different ways in one or several steps. It is known as

- A. Henry's law
- B. Joule's principle
- C. Hess's law
- D. Law of conservation of energy

Q.17 Solutions containing chlorate (I) ions are used as household bleaches and disinfectants. These solutions decompose on heating as shown  $3\text{ClO}^- \rightarrow \text{ClO}_3^- + 2\text{Cl}^-$

Which of the following ion will have highest oxidation state of chlorine

- A.  $\text{ClO}^-$
- B.  $\text{ClO}_2^-$
- C.  $\text{ClO}_3^-$
- D.  $\text{Cl}^-$

Q.18 Which of the following is a list of metals in order from strongest to weakest reducing agents

- A.  $\text{Au} > \text{Ni} > \text{Ag}$
- B.  $\text{Ni} > \text{Au} > \text{Ag}$
- C.  $\text{Ni} > \text{Ag} > \text{Au}$
- D.  $\text{Ag} > \text{Ni} > \text{Au}$

Q.19 Which of the following will be collected at cathode during electrolysis of aq.  $\text{CuSO}_4$  solution

- A.  $\text{H}_2$
- B.  $\text{O}_2$
- C.  $\text{SO}_3$
- D. Cu

Q.20 Enthalpy of combustion for carbon is  $-393.5 \text{ kJ/mole}$ . What is the enthalpy of formation of  $\text{CO}_2$  would be

- A.  $+393.5 \text{ kJmol}^{-1}$
- B.  $-393.5 \text{ kJmol}^{-1}$
- C.  $1 \text{ kJmol}^{-1}$
- D.  $0 \text{ kJmol}^{-1}$

Q.21 In exothermic reactions, the heat contents of the

- A. Products is more than that of reactants
- B. Reactants is equal to that of products
- C. Reactants is more than that of products
- D. Reactants is equal to zero

Q.22 If one mole of ammonia and one mole of hydrogen chloride are mixed in a closed container to form ammonium chloride gas, then

- A.  $\Delta H < \Delta E$
- B.  $\Delta H = \Delta E$
- C.  $\Delta H > \Delta E$
- D. No Relationship

Q.23 The conditions for standard enthalpy change are

- A. 760torr and 300K
- B. 760torr and 298K
- C. 760 torr and 277K
- D. 760torr and 0K

Q.24 Which of the following gives same products at cathode and anode during electrolysis as like water

- A.  $\text{CuSO}_4(\text{molten})$
- B.  $\text{CuSO}_4(\text{aq})$
- C.  $\text{Na}_2\text{SO}_4(\text{aq})$
- D.  $\text{NaCl}(\text{aq})$

Q.25 Stronger the reducing agent greater is the

- A. Oxidation potential
- B. Redox potential
- C. emf of cell
- D. Reduction potential

Q.26 Which of the following is incorrect statement about electrolytic cell

- A. Anode is positively charged
- B. Cathode is negatively charged
- C. Chemical energy is converted to electrical energy
- D. Reaction is non-spontaneous in electrolytic cell

Q.27  $E^\circ_{\text{red}}$  of an element can be calculated by comparing it with

- A. New electrode of same element
- B. 1M solution of ions of respective element
- C. Standard Hydrogen Electrode
- D. 2M solution of HCl

- Q.28  $O_{2(g)} + 2H_2O_{(l)} + xe^- \longrightarrow 4OH^-_{(aq)}$   
 The 'x' electrons in above equation are  
 A. 1  
 B. 3  
 C. 2  
 D. 4
- Q.29 One Joule is equal to  
 A. 4.184 Cal  
 B. 1/4.184 Cal  
 C. 7.60 Cal  
 D. 1.055 Cal
- Q.30 Consider the following reaction  
 $MnO_2 + HCl \longrightarrow MnCl_2 + H_2O + Cl_2$   
 Which atom undergoes decrease in oxidation number  
 A. Chlorine  
 B. Hydrogen  
 C. Oxygen  
 D. Manganese
- Q.31 In a reaction, the oxidation number of Cu decreases by 2. This indicates that Cu is  
 A. Reduced  
 B. Neutralized  
 C. Oxidized  
 D. A reducing agent.
- Q.32 When a Zn piece is placed in  $CuSO_4$  solution, copper gets precipitated because  
 A. Standard reduction potential of Zn is more than that of copper  
 B. Standard reduction potential of Zn is less than that of copper  
 C. Atomic number of Zn is larger than that of Cu  
 D. Atomic number of Zn is smaller than that of Cu
- Q.33 If the internal energy of the system is increased  
 A. Change in state of the system is increased  
 B. The temperature of the system may increase  
 C. The chemical reaction may take place  
 D. All
- Q.34 During balancing of Redox equations by oxidation number method the 'O' and H-atoms are balanced by  
 A. Method of difference  
 B. Inspection method  
 C. Ion-electron method  
 D. Spectroscopic method
- Q.35 Which of the following is incorrect about enthalpy of neutralization?  
 A. Its maximum value is 57.5kJ/mol  
 B. It is equal to enthalpy of formation of water from ions  
 C. It is more than 57.5kJ/mol when one of the acid or base is weaker  
 D. Energy is released in this process
- Q.36 Standard enthalpy of  $Al_2O_3$  cannot be measured because  
 A. It does not catch fire  
 B. It reacts with  $CO_2$   
 C. Protective layer of oxide covers the surface  
 D. It requires catalyst
- Q.37 The dissolution of ammonium chloride in water is a/an  
 A. Endothermic process  
 B. Exothermic process  
 C. Spontaneous process  
 D. Both "A" and "C"
- Q.38 In the reaction  $4Al + 3O_2 \rightarrow 4Al^{3+} + 6O^{2-}$ . Which of the following statements is incorrect  
 A. Metallic Aluminium is reducing agent  
 B.  $Al^{3+}$  is an oxidizing agent  
 C. Metallic Aluminium is reduced to  $Al^{3+}$   
 D. Redox reaction
- Q.39 If strip of Cu metal is placed in solution of  $FeSO_4$   
 A. Cu will be precipitated out  
 B. Fe will be precipitated out  
 C. Cu and Fe both dissolve  
 D. No reaction takes place
- Q.40 The process which is exothermic in nature  
 A. Atomization  
 B. Electrolysis  
 C. Hybridization  
 D. Ionization
- Q.41 According to first law of thermodynamics energy from system to surrounding can be transferred in the form of \_\_\_\_\_  
 A. Heat and work  
 B. Heat and force  
 C. Kinetic energy and pressure  
 D. Displacement and heat
- Q.42 When a solid melts there is  
 A. Decrease in enthalpy  
 B. Increase in enthalpy  
 C. No change in enthalpy  
 D. A decrease in internal energy
- Q.43 Boiling of water is \_\_\_\_\_ change  
 A. Endothermic  
 B. Non-energetic  
 C. Exothermic  
 D. Non-spontaneous
- Q.44 The expression  $\Delta E = q + w$  is  
 A. First law of the thermodynamics  
 B. Third law of the thermodynamics  
 C. Second law of the thermodynamics  
 D. Hess's Law



- Q.45 In acidic medium, moles of  $\text{H}_2\text{O}$  and  $\text{H}^+$  are added respectively for given reaction are  $10\text{Cl}^- + 2\text{MnO}_4^- \longrightarrow 2\text{Mn}^{2+} + 5\text{Cl}_2$
- A. 4, 4  
B. 8, 16  
C. 8, 8  
D. 4, 8
- Q.46 An enthalpy change which always proceeds through evolution of heat
- A.  $\Delta H_{\text{a}}$   
B.  $\Delta H_{\text{solution}}^{\circ}$   
C.  $\Delta H_{\text{n}}^{\circ}$   
D.  $\Delta H_{\text{f}}^{\circ}$
- Q.47 All of the following proceed through the oxidation except
- A.  $2\text{SO}_2 + \text{O}_2 \longrightarrow 2\text{SO}_3$   
B.  $\text{H}_2\text{SO}_3 \longrightarrow \text{H}_2\text{SO}_4$   
C.  $2\text{FeCl}_2 + \text{Cl}_2 \longrightarrow 2\text{FeCl}_3$   
D.  $\text{HNO}_3 \longrightarrow \text{HNO}_2$
- Q.48 The element having higher standard reduction potential than hydrogen
- A. Mg  
B. Au  
C. Sn  
D. Zn
- Q.49  $\text{C}_2\text{O}_4^{2-} \longrightarrow 2\text{CO}_2 + x\text{e}^-$   
The 'x' electrons in above equation are
- A. 1  
B. 3  
C. 2  
D. 4
- Q.50  $\text{NH}_4\text{OH} + \text{HCl} \longrightarrow \text{NH}_4\text{Cl} + \text{H}_2\text{O}$   
The enthalpy of above reaction will be
- A.  $-53 \text{ kJmol}^{-1}$   
B.  $+57.41 \text{ kJmol}^{-1}$   
C.  $-57.41 \text{ kJmol}^{-1}$   
D.  $-60 \text{ kJmol}^{-1}$
- Q.51 Which is used to measure enthalpy of combustion
- A. Glass calorimeter  
B. Bomb calorimeter  
C. Copper calorimeter  
D. Flame calorimeter
- Q.52 If number of moles of product is more than that of reactant than in gaseous reaction
- A.  $q_p = q_v$   
B.  $q_p > q_v$   
C.  $q_p < q_v$   
D.  $q_p = q_v/2$
- Q.53 The amount of heat required to convert one mole of a solid directly into its vapour state at STP is called as
- A. Standard heat of sublimation  
B. Heat of reaction  
C. Molar heat of vaporization  
D. Heat of neutralization
- Q.54 The system becomes less stable if it
- A. Gains energy  
B. Does not lose energy  
C. Loses energy  
D. Have constant energy
- Q.55 The total energy of the system is equal to
- A. P.E + K.E  
B. P.E + heat energy  
C. K.E + heat energy  
D. P.E + mechanical energy
- Q.56 Which of the following is correct representation for calculation of lattice energy?
- A.  $\Delta H_{\text{Latt.}} = \Delta H_{\text{f}} - \Delta H_{\text{x}}$   
B.  $\Delta H_{\text{Latt.}} = \Delta H_{\text{f}} + \Delta H_{\text{x}}$   
C.  $\Delta H_{\text{x}} = \Delta H_{\text{f}} + \Delta H_{\text{Latt.}}$   
D.  $\Delta H_{\text{Latt.}} = \Delta H_{\text{x}} - \Delta H_{\text{f}}$
- Q.57 Standard reduction electrode potentials of three metals P, Q and R are +0.5 V, -3.0 V and -1.2 V respectively. The order of reducing power of these metals is
- A.  $Q > R > P$   
B.  $R > Q > P$   
C.  $P > Q > R$   
D.  $P > R > Q$
- Q.58 When  $\text{HNO}_3$  acts as an oxidizing agent and ultimately forms  $\text{NO}_2$ ,  $\text{NO}$ ,  $\text{HNO}_2$  then the number of electrons transferred in each case is
- A. 1, 3, 2  
B. 4, 3, 1  
C. 1, 5, 3  
D. 3, 5, 7
- Q.59 Through the salt bridge there is a transfer of
- A. Current  
B. Ions  
C. Free Radicals  
D. Molecules
- Q.60 Among the following molecules, in which does Phosphorous show the lowest oxidation number
- A.  $\text{HPO}_3$   
B.  $\text{PCl}_3$   
C.  $\text{PH}_3$   
D.  $\text{P}_4$

| A  | B | C | D | A | B  | C | D | A | B | C  | D | A | B | C | D  |   |   |   |   |
|----|---|---|---|---|----|---|---|---|---|----|---|---|---|---|----|---|---|---|---|
| 1  | ○ | ● | ● | ○ | 16 | ○ | ○ | ● | ○ | 31 | ● | ○ | ○ | ○ | 46 | ○ | ○ | ○ | ● |
| 2  | ○ | ○ | ○ | ● | 17 | ○ | ○ | ○ | ○ | 32 | ○ | ● | ○ | ○ | 47 | ○ | ○ | ○ | ● |
| 3  | ○ | ○ | ● | ○ | 18 | ○ | ○ | ○ | ○ | 33 | ○ | ○ | ○ | ● | 48 | ○ | ● | ○ | ○ |
| 4  | ○ | ○ | ○ | ○ | 19 | ○ | ○ | ○ | ○ | 34 | ○ | ○ | ○ | ○ | 49 | ○ | ○ | ○ | ○ |
| 5  | ○ | ○ | ○ | ○ | 20 | ○ | ○ | ○ | ○ | 35 | ○ | ○ | ○ | ○ | 50 | ○ | ○ | ○ | ○ |
| 6  | ○ | ○ | ○ | ○ | 21 | ○ | ○ | ○ | ○ | 36 | ○ | ○ | ○ | ○ | 51 | ○ | ○ | ○ | ○ |
| 7  | ○ | ○ | ○ | ○ | 22 | ○ | ○ | ○ | ○ | 37 | ○ | ○ | ○ | ○ | 52 | ○ | ○ | ○ | ○ |
| 8  | ○ | ○ | ○ | ○ | 23 | ○ | ○ | ○ | ○ | 38 | ○ | ○ | ○ | ○ | 53 | ○ | ○ | ○ | ○ |
| 9  | ○ | ○ | ○ | ○ | 24 | ○ | ○ | ○ | ○ | 39 | ○ | ○ | ○ | ○ | 54 | ○ | ○ | ○ | ○ |
| 10 | ○ | ○ | ○ | ○ | 25 | ○ | ○ | ○ | ○ | 40 | ○ | ○ | ○ | ○ | 55 | ○ | ○ | ○ | ○ |
| 11 | ○ | ○ | ○ | ○ | 26 | ○ | ○ | ○ | ○ | 41 | ○ | ○ | ○ | ○ | 56 | ○ | ○ | ○ | ○ |
| 12 | ○ | ○ | ○ | ○ | 27 | ○ | ○ | ○ | ○ | 42 | ○ | ○ | ○ | ○ | 57 | ○ | ○ | ○ | ○ |
| 13 | ○ | ○ | ○ | ○ | 28 | ○ | ○ | ○ | ○ | 43 | ○ | ○ | ○ | ○ | 58 | ○ | ○ | ○ | ○ |
| 14 | ○ | ○ | ○ | ○ | 29 | ○ | ○ | ○ | ○ | 44 | ○ | ○ | ○ | ○ | 59 | ○ | ○ | ○ | ○ |
| 15 | ○ | ○ | ○ | ○ | 30 | ○ | ○ | ○ | ○ | 45 | ○ | ○ | ○ | ○ | 60 | ○ | ○ | ○ | ○ |

| Roll No. |   |   |   |   |   |   |
|----------|---|---|---|---|---|---|
|          |   |   |   | 2 | 6 | 5 |
| 0        | 0 | 0 | 0 | 0 | 0 | 0 |
| 1        | 1 | 1 | 1 | 1 | 1 | 1 |
| 2        | 2 | 2 | 2 | 2 | 2 | 2 |
| 3        | 3 | 3 | 3 | 3 | 3 | 3 |
| 4        | 4 | 4 | 4 | 4 | 4 | 4 |
| 5        | 5 | 5 | 5 | 5 | 5 | 5 |
| 6        | 6 | 6 | 6 | 6 | 6 | 6 |
| 7        | 7 | 7 | 7 | 7 | 7 | 7 |
| 8        | 8 | 8 | 8 | 8 | 8 | 8 |
| 9        | 9 | 9 | 9 | 9 | 9 | 9 |

1. Use Blue Ball
  2. Please Fill In T
  3. It Is Important Completely And In The Examp
- Correct Example
- Incorrect Example

SUBJECT: Chem CH/UNIT#    

NAME: MHT

R. NO. (IN WORDS): 265

CLASS: S-4 SESSION:    

SIGN: [Signature] DATE:

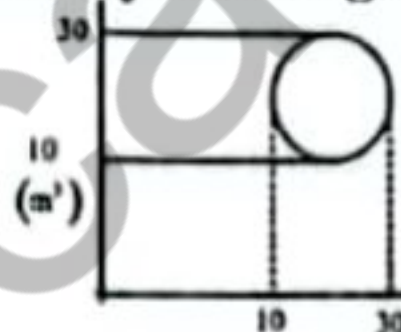
# PHYSICS MDCAT

## UNIT-5 (A+ SERIES)

### TOPICS:

#### ✓ Thermodynamics

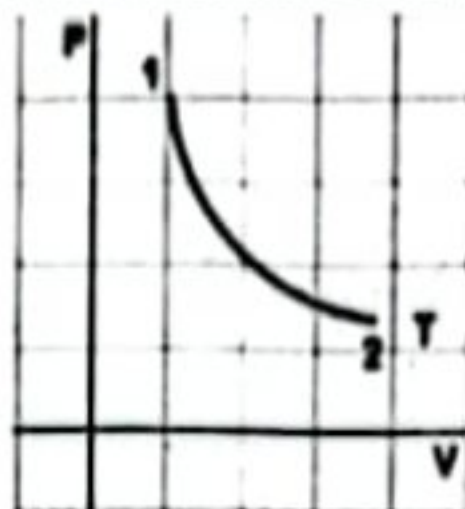
- Q.1 If 20J of work is done in compressing a gas adiabatically the change in internal energy is equal to  
A. 20J  
B. 10J  
C. -20J  
D. 200J
- Q.2 The direction of flow of heat between two bodies is determined by  
A. Internal energy  
B. Total energy  
C. Kinetic energy  
D. None of these
- Q.3 A cycle tyre bursts suddenly. This represents an  
A. Isothermal process  
B. Isochoric process  
C. Isobaric process  
D. Adiabatic process
- Q.4 If 315cal of heat is given to the system, and the system does 20cal of work, find the change in internal energy.  
A. 335cal  
B. 295cal  
C. 335J  
D. 0 J
- Q.5 Suppose volume of gas in a cylinder is  $3 \text{ cm}^3$ , if the piston is kept fixed and gas is heated from  $5^\circ\text{C}$  to  $12^\circ\text{C}$  then the work done is  
A. 2.3 J  
B. 21 J  
C. Zero  
D. 25 J
- Q.6 Consider a gas contained in a rigid container of volume  $20\text{m}^3$ . What will be the change in internal energy if 50J of heat is provided to it? Assume the gas exerts a pressure of 1atm on the walls.  
A. 50J  
B. 0  
C. 70J  
D. 20J
- Q.7 When two bodies are said to be in thermal equilibrium, then net exchange of heat between them is:  
A. Constant  
B. Zero  
C. Infinite  
D. Continuous flow
- Q.8 The change in internal energy by the system in going through cycle is



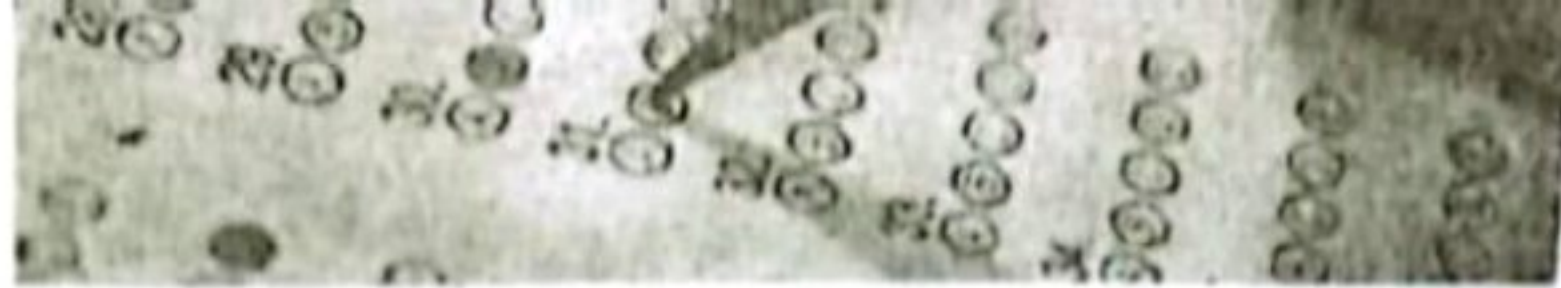
- A.  $10^7 \pi \text{ J}$   
B.  $10^{-3} \text{ J}$   
C. 0  
D.  $10^2 \pi \text{ J}$
- Q.9 Which one of the following process has greater slope in P – V diagram; (Take, V, along x-axis and 'P' along y-axis)  
A. Isothermal process  
B. Process carried out at constant pressure  
C. Adiabatic process  
D. All have same slope
- Q.10 If  $C_p$  and  $C_v$  are the molar specific heats of a gas at constant pressure and volume respectively then the ratio of adiabatic and isothermal moduli of elasticity will be:  
A.  $\frac{C_p - C_v}{C_p}$   
B.  $\frac{C_v}{C_p}$   
C.  $C_p C_v$   
D.  $\frac{C_p}{C_v}$
- Q.11 A gas is being compressed adiabatically. The specific heat of the gas during compression is  
A. Zero  
B. Finite but non zero  
C. Infinite  
D. Undefined
- Q.12 A system undergoes an adiabatic process in which its internal energy increases by 20 J. Which of the following statements is true?  
A. 20 J of work was done on the system  
B. The system received 20 J of energy as heat  
C. 20 J of work was done by the system  
D. The system lost 20 J of energy as heat
- Q.13 For the Boyle's law to hold good, the necessary condition is:  
A. Isothermal  
B. Isobaric  
C. Adiabatic  
D. Isochoric



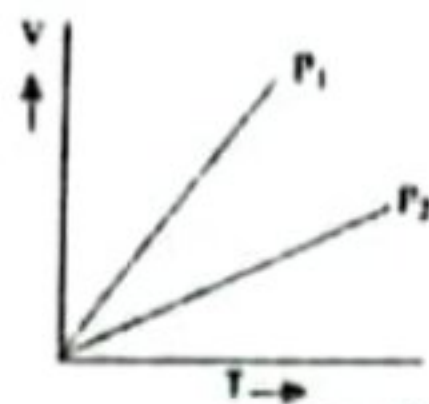
- Q.14 For hydrogen gas  $C_p - C_v = a$ , and for oxygen gas  $C_p - C_v = b$ , so that relation between a and b given by  
 A.  $a = 16b$   
 B.  $a = b$   
 C.  $16a = b$   
 D.  $a = 4b$
- Q.15 The state of an ideal gas is changed isothermally from position 1 to position 2 is shown below. What is the change in the internal energy of the gas during this process?



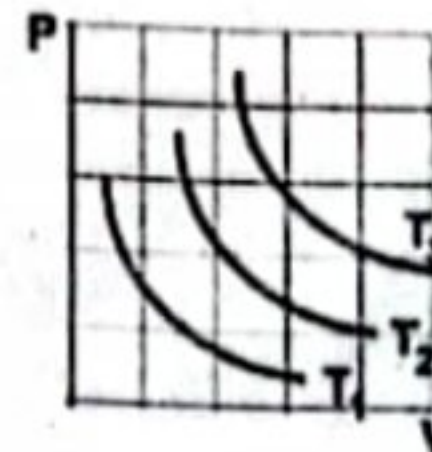
- A.  $\Delta U = W$   
 B.  $\Delta U = Q$   
 C.  $\Delta U > 0$   
 D.  $\Delta U = 0$
- Q.16 The graph of isochoric process is called  
 A. Adiabatic  
 B. Isotherm  
 C. Isochor  
 D. Isobar
- Q.17  $C_m =$  (where  $C_m$  is constant of proportionality and is known as molar specific heat)  
 A.  $\frac{n\Delta Q}{\Delta T}$   
 B.  $\frac{\Delta Q}{n\Delta T}$   
 C.  $\frac{\Delta Q}{n^2\Delta T}$   
 D.  $\frac{\Delta Q}{n^3\Delta T}$
- Q.18 200 g water is heated from  $40^\circ\text{C}$  to  $60^\circ\text{C}$ . Ignoring the slight expansion of water, the change in its internal energy is close to (Given specific heat of water =  $4184\text{ J/kgK}$ )  
 A. 167.4 kJ  
 B. 8.4 kJ  
 C. 4.2 kJ  
 D. 16.7 kJ
- Q.19 The internal energy of the gas increases in  
 A. Adiabatic expansion  
 B. Adiabatic compression  
 C. Isothermal expansion  
 D. Isothermal compression
- Q.20 An ideal monatomic gas at constant volume has a molar specific heat  $C_v$ :  
 A.  $R$   
 B.  $3R/2$   
 C.  $5R/2$   
 D.  $7R/2$
- Q.21 How much heat is needed to raise the temperature of a block of copper (with mass 0.5 kg) from  $0^\circ\text{C}$  to  $100^\circ\text{C}$ ? (for copper,  $c = 386\text{ J/kg}^\circ\text{C}$ )  
 A. 21 kJ  
 B. 19.3 kJ  
 C. 12.4 kJ  
 D. 6.25 kJ
- Q.22 When 6400 J of heat is supplied to the body and it raises its temperature by approximately 100 degrees, then what is its heat capacity?  
 A. 64 J/K  
 B. 98 J/K  
 C. 46 J/K  
 D. 28 J/K
- Q.23 What is not true for isothermal process  
 A.  $\Delta T > 0$   
 B.  $\Delta U = 0$   
 C.  $\Delta Q = \Delta W$   
 D.  $PV = \text{constant}$
- Q.24 If  $P = P_0$  and  $V = V_0$  gas expands isothermally to  $P = 3P_0$  then Volume is  
 A.  $3V_0$   
 B.  $2V_0$   
 C.  $V_0/3$   
 D.  $V_0/2$
- Q.25 If the system goes from two different paths to same final state, then  $Q_1$  and  $W_1$  and  $Q_2$  and  $W_2$  are heat absorbed and work done then  
 A.  $Q_1 = Q_2$   
 B.  $W_1 = W_2$   
 C.  $Q_1 + W_1 = Q_2 + W_2$   
 D.  $Q_1 - W_1 = Q_2 - W_2$
- Q.26 Value of  $C_p - C_v$  for a monoatomic gas is:  
 A. 1.67  
 B. 1.40  
 C. 1.29  
 D. None of these
- Q.27 Units of molar specific heat are same as that of  
 A. Pressure  
 B. Entropy  
 C. Molar gas constant  
 D. Temperature
- Q.28 Heat capacity is  
 A. Sum of mass of object and specific heat  
 B. Difference of mass and specific heat  
 C. Product of mass and specific heat  
 D. None
- Q.29 Addition of heat at constant pressure to give a result in  
 A. Raising its temperature  
 B. Raising its pressure  
 C. Raising its volume  
 D. Raising temperature and doing external work



- Q.30 When the system is expanded by adding heat energy then the work done is:  
 A. Positive and on the system  
 B. Negative and on the system  
 C. Positive and by the system  
 D. Negative and by the system
- Q.31 Which of the terms is related with thermodynamics?  
 A. System  
 B. Surrounding  
 C. Boundary  
 D. All of these
- Q.32 For an adiabatic expansion  
 A.  $\Delta U = -ve$   
 B.  $W = -ve$   
 C.  $\Delta U = 0$   
 D.  $\Delta T = 0$
- Q.33 Substance which has maximum specific heat capacity is  
 A. Copper  
 B. Water  
 C. Aluminium  
 D. Mercury
- Q.34 Graph between pressure and temperature for an ideal gas at constant volume is  
 A. Straight line  
 B. Ellipse  
 C. Parabola  
 D. Sinusoidal
- Q.35 In which terms expressed hotness or coldness of an object:  
 A. Temperature  
 B. Heat  
 C. Chemical energy  
 D. Thermal energy
- Q.36 For an ideal gas V-T curves as constant pressures  $P_1$  &  $P_2$  are shown in figure - from the figure



- A.  $P_1 > P_2$   
 B.  $P_1 < P_2$   
 C.  $P_1 = P_2$   
 D.  $P_1 \geq P_2$
- Q.37 What are the factor/factors on which the quantity of heat given to a body depends?  
 A. The mass of the body  
 B. The rise (or fall) in the temperature of the body  
 C. Nature of the material of the body  
 D. All of these
- Q.38 Energy supplied to convert unit mass of substance from solid to liquid state at its melting point is called  
 A. Latent heat of fusion  
 B. Evaporation  
 C. Solidification  
 D. Latent heat of fission
- Q.39 The amount of heat required to raise the temperature of 10 moles of water from 70K to 80K (molar heat capacity of water 75.24J) is:  
 A. 0.7524J  
 B. 7524J  
 C. 95.24J  
 D. 572.4J
- Q.40 First law of thermodynamics is applicable for  
 A. Electric potential energy  
 B. Heat energy  
 C. Strain Energy  
 D. Sound energy
- Q.41 The internal energy of piece of lead when beaten by a hammer will.  
 A. Increase  
 B. Remain constant  
 C. Decrease  
 D. Sometimes increase and sometime decrease
- Q.42 The state of an ideal gas was changed three times at three different temperatures. The diagram represents three different isothermal curves. Which of the following is true about the temperature of the gas?  
 A.  $T_1 > T_2 > T_3$   
 B.  $T_1 < T_2 < T_3$   
 C.  $T_1 > T_2 < T_3$   
 D.  $T_1 > T_2 = T_3$



- Q.43 The energy transforming processes that occur within an organism are named as  
 A. Metabolism  
 B. Food chain  
 C. Catabolism  
 D. Mitochondria
- Q.44 50 J heat is transferred slowly to a gas which pushes the piston of cross-sectional area  $0.2 \text{ m}^2$  through a distance of 2.0 cm and pressure of the gas is maintained at  $10000 \text{ Nm}^{-2}$ . Change in internal energy of gas during expansion is  
 A. 10 J  
 B. -10 J  
 C. 20 J  
 D. -20 J
- Q.45 Which of the following parameters does not characterize the thermodynamic state of matter?  
 A. Temperature  
 B. Pressure  
 C. Work  
 D. Volume
- Q.46 A system is given 400 cal of heat and 1000 J of work is done by the system, then the change in internal energy of the system will be.  
 A. -860 J  
 B. 680 erg  
 C. 680 J  
 D. 660 J

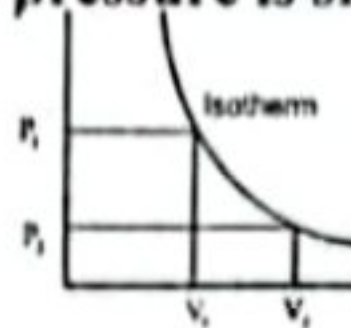


- Q.47 An adiabatic process is also called  
 A. Heat exchange process  
 B. Heating process  
 C. Isentropic process  
 D. All of the above
- Q.48 The melting point of copper is  $1,080^{\circ}\text{C}$  and its heat of fusion is  $200 \text{ kJ/kg}$ . If a copper coin at this temperature is completely melted by the absorption of  $2,000 \text{ J}$  of heat, what is the mass of the coin?  
 A.  $\frac{1}{1,080} \text{ kg}$   
 B.  $\frac{1}{540} \text{ kg}$   
 C.  $\frac{1}{108} \text{ kg}$   
 D.  $\frac{1}{100} \text{ kg}$

- Q.49 When the temperature of the gas system is changed by  $30\text{K}$  and the internal energy of the system is changed by  $48\text{J}$ , the molar specific heat of system at constant volume will be:  
 A.  $0.6\text{J/mol.K}$   
 B.  $1.0\text{J/mol.K}$   
 C.  $1.6\text{J/mol.K}$   
 D.  $2.1\text{J/mol.K}$

- Q.50 The volume of  $1\text{m}^3$  of gas is doubled at atmospheric pressure. The work done at constant pressure will be  
 A. 0  
 B.  $10^5 \text{ cal}$   
 C.  $10^5 \text{ J}$   
 D.  $10^5 \text{ erg}$

- Q.51 What will be the mathematical form of first law of thermodynamics for a system whose variation of volume with pressure is shown?



- A.  $Q = U$   
 B.  $Q = \frac{U}{W}$   
 C.  $U = W$   
 D.  $Q = W$
- Q.52 If the specific heat capacity of iron is  $0.108 \text{ cal/g}^{\circ}\text{C}$ . Then how much heat is sufficient to raise the temperature of  $1\text{g}$  of iron through  $1^{\circ}\text{C}$ .  
 A.  $0.108 \text{ calories}$   
 B.  $108 \text{ calories}$   
 C.  $0.0108 \text{ calories}$   
 D.  $10.8 \text{ calories}$

- Q.53 In an isothermal process the volume of an ideal gas is halved. One can say that  
 A. Internal energy of the system decreases  
 B. Work done by the gas is positive  
 C. Work done by the gas is negative  
 D. Internal energy of the system increases

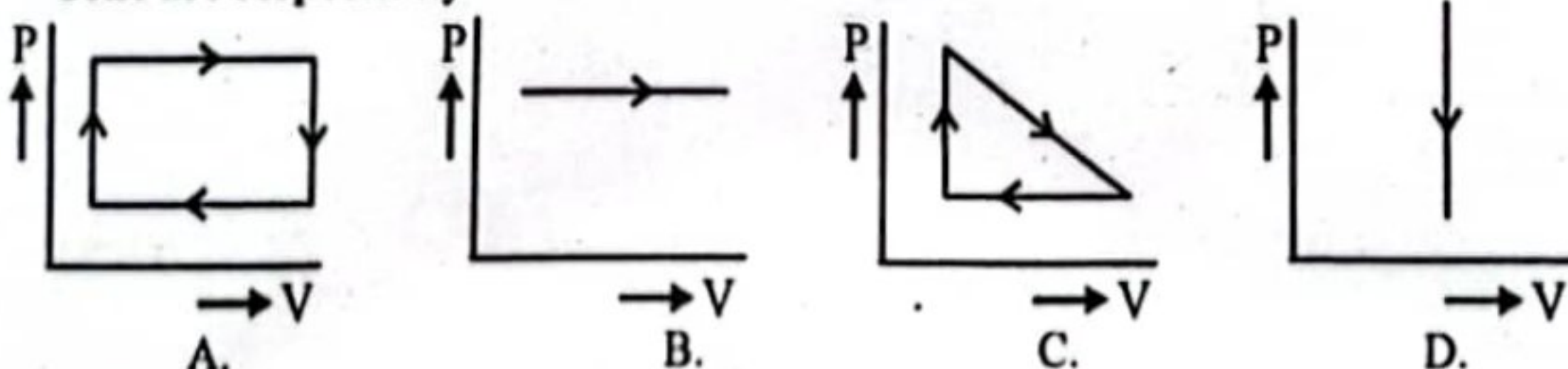
- Q.54 In which of the following processes all three thermodynamic variables, that is, pressure, volume and temperature can change?

- A. Isobaric  
 B. Isothermal  
 C. Isochoric  
 D. Adiabatic
- Q.55 If  $C_p = 34.4 \text{ J/mol K}$  and value of  $R$  is  $8.314$  then  
 A.  $C_v = 26.1 \text{ J/mol K}$   
 B.  $C_v = 20.1 \text{ J/mol K}$   
 C.  $C_v = 10.1 \text{ J/mol K}$   
 D.  $C_v = 120.1 \text{ J/mol K}$

- Q.56 Heat added at constant volume of a gas is used to  
 A. To do external work  
 B. To increase its internal energy  
 C. Either "A" or "C"  
 D. Both "A" and "C"

- Q.57 When a substance is heated, energy associated with its atoms or molecules:  
 A. Increases  
 B. Remains same  
 C. Decreases  
 D. Becomes zero

- Q.58 The indicator diagrams representing maximum and minimum amounts of work done are respectively



- A. (A) and (C)  
 B. (C) and (D)  
 C. (C) and (B)  
 D. (B) and (D)
- Q.59 Work done by the system on its environment is taken as:  
 A. Positive  
 B. Negative  
 C. Neutral  
 D. None of these
- Q.60 The K.E of molecules of an ideal gas at absolute zero will be:  
 A. Zero  
 B. Infinite  
 C. Very high  
 D. Below zero

| A  | B                                | C                                | D                                | A                                | B  | C                                | D                                | A                                | B                                | C  | D                                | A                                | B                                | C                                | D  |                                  |                                  |                                  |                                  |
|----|----------------------------------|----------------------------------|----------------------------------|----------------------------------|----|----------------------------------|----------------------------------|----------------------------------|----------------------------------|----|----------------------------------|----------------------------------|----------------------------------|----------------------------------|----|----------------------------------|----------------------------------|----------------------------------|----------------------------------|
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**Roll No.**

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| 0 | 0 | 0 | 0 | 0 | 0 | 0 |
| 1 | 1 | 1 | 1 | 1 | 1 | 1 |
| 2 | 2 | 2 | 2 | 2 | 2 | 2 |
| 3 | 3 | 3 | 3 | 3 | 3 | 3 |
| 4 | 4 | 4 | 4 | 4 | 4 | 4 |
| 5 | 5 | 5 | 5 | 5 | 5 | 5 |
| 6 | 6 | 6 | 6 | 6 | 6 | 6 |
| 7 | 7 | 7 | 7 | 7 | 7 | 7 |
| 8 | 8 | 8 | 8 | 8 | 8 | 8 |
| 9 | 9 | 9 | 9 | 9 | 9 | 9 |

**SUBJECT:** Phy

**NAME:** M. A.

**R. NO. (IN WORDS):** \_\_\_\_\_

**CLASS:** S-4

**SIGN:** H-107

# ENGLISH MDCAT

## TEST-5 (A+ SERIES)

### TOPICS:

#### MISCELLANEOUS MISTAKES

#### MODIFIERS

#### Directions:

Spot the Error.

- Q.1 As (A) the evening advanced ahead(B), the cold became more intense(C), and it induced people to seek comfort in pleasure(D)  
A. As B. advanced ahead C. more intense D. in pleasure.
- Q.2 I do not know that what(A) you did all this for(B), and I am not going(C) to ask you to give me any justification.(D)  
A. that what B. all this for C. am not going D. any justification
- Q.3 George, along with(A) other newcomers, has not completed(B) the assignment yet and(C) Marie hasn't neither.(D)  
A. along with B. has not completed C. yet and D. hasn't neither
- Q.4 The underlying (A) motive behind (A) his kind words was so that(A) he might borrow (D) money from me.  
A. underlying B. behind C. so that D. might borrow
- Q.5 Modern methods of selecting(A) and to process(B) foods have removed(C) the risk of poisoning from canned foods.(D)  
A. selecting B. to process C. have removed D. from canned food
- Q.6 The Magi (A), who(B) brought valuable gifts(C) to newborn Christ, they(D) were very wise people.  
A. The Magi B. who C. valuable gifts D. they
- Q.7 On the Martian hills, the people from the Earth saw a few(A) people who(B) were friendly(C) and nicely.(D)  
A. a few B. who C. friendly D. nicely
- Q.8 It(A) is a crisis(B) for which(C) she was unprepared for.(D)  
A. It B. a crisis C. which D. for
- Q.9 Suddenly(A) he stopped at the edge of the meadow(B), taking his pocket knife(C) from his pocket, and cut a wisp of alfalfa.(D)  
A. Suddenly B. stopped at the edge of the meadow  
C. taking his knife D. cut a wisp of alfalfa
- Q.10 The tags remain active(A) after we have left the store unless(B) they (C) are not switched off.(D)  
A. remain active B. unless C. they D. not switched off
- Q.11 The sun and the other stars (A) we see in the sky they (B) are all too hot for life to exist.  
A. the other stars B. they (remove they) C. for D. to exist.
- Q.12 Humanitarian disasters (A), of which genocide is the most(B) appalling of them (C), are not pretty (D) things.  
A. disasters B. most C. of them D. pretty.
- Q.13 The results(A) of the study were inconclusive (B) therefore(C) more research needs to be done (D) on the topic.  
A. the results B. inconclusive, C. therefore D. to be done
- Q.14 During (A) the inquisition, Jack was asked that (B) if he had (C) had other (D) affairs.  
A. during B. that C. had D. other
- Q.15 We ordered a more bigger (A) cake than (B) usual in case unexpected (C) guests came to (D) the party.  
A. more bigger B. than C. unexpected D. to
- Q.16 Chips could not (A) scarcely walk, and it (B) was a hard job getting (C) him down (D) the track.  
A. not B. it C. getting D. down
- Q.17 He would then retreat back, (A) hoping to catch (B) the elusive (C) man in their (D) new round.  
A. retreat back B. to catch C. elusive D. their
- Q.18 You can find these (A) books either (B) on (C) the table or the shelf. (D)  
A. these B. either C. on D. the shelf
- Q.19 The guru wants his students not only (A) to keep (B) quiet but also(C) doing (D) the task.  
A. not only B. to keep C. but also D. doing
- Q.20 It was a real fact (A) that its discovery has stimulated (B) new research (C) to find something (D) better.  
A. real fact B. has stimulated C. new research D. something

#### Directions:

Choose the Correct Option

Q.21

- A. At five years old, my father took me to get a haircut for the very first time.  
B. At five years old, my father took me to get my first haircut.  
C. My father took me at five years old to get a haircut for the very first time.  
D. When I was five, my father took me for my first haircut.

Q.22

- A. She bought a long beautiful white French silk wedding dress.  
B. She bought a beautiful long white French wedding silk dress.  
C. She bought a beautiful white long French silk wedding dress.  
D. She bought a beautiful long white French silk wedding dress.



- Q.23 A. I have watched a lot of interesting movies lately. B. I have watched a lot of interested movies lately.  
C. I have watched a lot of interesting movies latest. D. I have watched a lot of interested movies recently.
- Q.24 A. If you don't have money enough, I can lend you some.  
B. If you don't have enough money, I can lend you some.  
C. If you don't have enough money, I could lend you some.  
D. If you don't have enough money, I can borrow you some.
- Q.25 A. My new computer is much better than my old one. B. My new computer is very better than my old one.  
C. My new computer is much better than me old one. D. My new computer is too better than that of my old.
- Q.26 A. The valley lay quietly and peacefully in the sun.  
B. The valley lay in a quietly and peacefully way in the sun.  
C. The valley lay in quiet and peaceful in the sun.  
D. The valley lay quiet and peaceful in the sun.
- Q.27 A. Dad walks impatiently into town every afternoon before supper to get a newspaper.  
B. Dad walks into town impatiently every afternoon before supper to get a newspaper.  
C. Dad walks every afternoon impatiently into town before supper to get a newspaper.  
D. Dad impatiently walks into town every afternoon before supper to get a newspaper.
- Q.28 A. The work was hard at the beginning, but she early got used to it.  
B. The work was hard at the beginning, but she soon got used to it.  
C. The work was hardly at the beginning, but she soon got used to it.  
D. The work was hardly at the beginning, but she early got use to it.
- Q.29 A. Whenever Mr. Steward came to meet Norma, she smiled friendly.  
B. Whenever Mr. Steward came to meet Norma, she smiled in a friendly manner.  
C. Whenever Mr. Steward came to meet Norma, she smiled friendly manner.  
D. Whenever Mr. Steward came to meet Norma, she smiled in a friendship manner.
- Q.30 A. The new President was welcomed in the evening at the airport cordially.  
B. The new President was welcomed at the airport cordially in the evening.  
C. The new President was cordially welcomed in the evening at the airport.  
D. The new President was welcomed cordially at the airport in the evening.

**Directions:**

**Fill in the Blanks**

- Q.31 She is the \_\_\_\_\_ of the two sisters.  
A. older B. oldest C. elder D. eldest
- Q.32 His wife's rude behaviour gives him \_\_\_\_\_ pain.  
A. much so B. much too C. much very D. too much
- Q.33 I shall be visiting \_\_\_\_\_.  
A. For a few days in Paris in July B. Paris in July for a few days  
C. in July Paris for a few days D. Paris for a few days in July
- Explanation:**  
(Place + Frequency + Time)
- Q.34 Our relations will be \_\_\_\_\_ with them.  
A. the most friendly B. a most friendly C. most friendly D. the most friendliest
- Q.35 The Martian people built a \_\_\_\_\_ cottage.  
A. white small concrete B. white concrete small C. concrete small white D. small white concrete
- Q.36 Aisha is \_\_\_\_\_ of the two candidates.  
A. better B. best C. the better D. the best
- Q.37 He is \_\_\_\_\_ me.  
A. more intelligent as B. the most intelligent as C. as intelligent as D. as more intelligent as
- Q.38 John's car wasn't so comfortable as \_\_\_\_\_.  
A. that of Ali's B. Ali's C. the car of Ali's D. Ali
- Q.39 \_\_\_\_\_, the trees looked like spooky aliens.  
A. while Walking home at night B. As Jonas was walking home at night  
C. when Walking home at night D. As Walking home at night
- Q.40 Painting for three hours at night, \_\_\_\_\_.  
A. Maggie finally finishing the kitchen B. the kitchen finally finished by Maggie  
C. the kitchen was finally finished by Maggie D. Maggie finally finished the kitchen

A B C D

1 ○ ● ○ ○

2 ● ○ ○ ○

3 ○ ○ ○ ●

4 ● ○ ● ○

5 ○ ● ○ ○

6 ○ ○ ○ ●

7 ○ ○ ○ ●

8 ○ ○ ○ ●

9 ○ ○ ● ○

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12 ○ ○ ● ○

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A B C D

16 ● ○ ○ ○

17 ● ○ ○ ○

18 ● ○ ○ ●

19 ○ ○ ○ ●

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A B C D

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42 ○ ○ ○ ○

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